Negative Ions
(Anions)

Monatomic

Rule:
Stem of the element name + ide
Examples:
H⁻ hydride ion
F⁻ fluoride ion
O²⁻ oxide ion
N³⁻ nitride ion

Oxyanions
(Containing Oxygen)

Rule:
Least oxygen: hypo_ite ion
Less oxygen: _ite ion
More oxygen: _ate ion
Most oxygen: per_ate ion
Examples:
ClO⁻ hypochlorite ion
ClO₂⁻ chlorite ion
ClO₃⁻ chlorate ion
ClO₄⁻ perchlorate ion
SO₃²⁻ sulfate ion
SO₄²⁻ sulfite ion

Comment: Halogens (except F) form all four ions. Most others only form two ions, the _ite and the _ate ions.

Oxyanions
(Containing Hydrogen)

Rule:
H oxyanion: “hydrogen” + name of oxyanion or “bi” + oxyanion
Examples:
HCO₃⁻ hydrogen carbonate ion or bicarbonate ion
HSO₄⁻ hydrogen sulfate ion or bisulfate ion
H₂PO₄⁻ dihydrogen phosphate ion

Comment:
H₂CO₃ is not named according to this rule because it is a compound, not an ion.

Others and Exceptions

Rule:
These items do not follow any rules; they must be memorized.
Examples:
OH⁻ hydroxide ion
CN⁻ cyanide ion
SCN⁻ thiocyanate ion
O₂²⁻ peroxide ion
C₂H₃O₂⁻ acetate ion